



## Subject card

Subject name and code	Inorganic Chemistry, PG_00048754						
Field of study	Green Technologies						
Date of commencement of studies	October 2020	Academic year of realisation of subject			2020/2021		
Education level	first-cycle studies	Subject group			Obligatory subject group in the field of study		
Mode of study	Full-time studies	Mode of delivery			at the university		
Year of study	1	Language of instruction			Polish		
Semester of study	1	ECTS credits			4.0		
Learning profile	general academic profile	Assessment form			assessment		
Conducting unit	Department of Inorganic Chemistry -> Faculty of Chemistry						
Name and surname of lecturer (lecturers)	Subject supervisor		dr hab. Katarzyna Kazimierczuk				
	Teachers		dr hab. Katarzyna Kazimierczuk dr inż. Daria Kowalkowska-Zedler				
Lesson types and methods of instruction	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	30.0	30.0	0.0	0.0	0.0	60
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	60		5.0		35.0	100
Subject objectives	The aim of the course is to teach students the direction of Zielone Technologie of the broadly understood basis of chemistry.						
Learning outcomes	Course outcome		Subject outcome		Method of verification		
	[K6_U05] can formulate and solve engineering tasks analytical methods, simulation as well as experimental, able to apply knowledge of basic physics and mathematics to analyze the results of experiments, is able to analyze and assess existing technical solutions		Student is able to apply the knowledge in the field of inorganic chemistry together with the laws describing the foundations of physics and mathematics.		[SU4] Assessment of ability to use methods and tools [SU2] Assessment of ability to analyse information		
	[K6_W02] has a basic knowledge of chemistry including general chemistry, inorganic, organic, physical, analytical, including the knowledge necessary to describe and understand the phenomena and chemical processes occurring in the environment; measurement and the determination of the parameters of these processes.		Student can invoke and apply the basic chemical laws and definitions of general, inorganic, physical, organic and analytical chemistry useful in analysis of chemical processes occurring in the environment.		[SW3] Assessment of knowledge contained in written work and projects [SW1] Assessment of factual knowledge		

Subject contents	<p>Lecture</p> <p>1. Basic concepts and chemical laws: What is chemistry? Chemical substance, elements and chemical compounds. Atom and molecule. Mole. Atomic mass. Molecular atomic and molecular mass. Law of conservation of mass and energy. The law of fixed and multiple relations. The law of simple volumetric relations. Chemical compound and mixture. Solutions. Ways of expressing composition and concentration.</p> <p>2. Chemical reactions: Chemical equations. Types of chemical reactions: synthesis, analysis, and exchange. Combustion reaction. The reactions occurring in solutions and precipitation reactions. Acid-base reactions. Oxidation and reduction reactions. Endo- and exothermic reactions. Photochemical reactions. Stoichiometry, nomenclature of chemical compounds: Nomenclature principles for basic inorganic compounds. Common and systematic names of hydrides, oxides, hydroxides, acids and salts. Nomenclature of selected groups of organic compounds. Structural and spatial isomerism. Isomerism of the position and isomerization of functional groups. Geometric isomerism and optical isomers. Empirical formula and molecular formula. Determination of molecular weight.</p> <p>3. Electronic structure of atom and periodic system: Quantization of energy. Absorption spectra and spectrum emission. Quantum Planck Condition. Hydrogen Bohr model. Dual-wave corpuscularism. Waves de Broglie. Spin electron. Heisenberg's uncertainty principle. Wave function and its physical meaning. Schrödinger equation and solution idea. Quantum numbers. Radial distribution function. Atomic orbitals. Principles of shell extension: Hund rule and Pauli rule. Hydrogen atom. Hydrogen-like atoms. Multi-electron atoms. Electron configurations of atoms. Periodic table. Periodicity of properties. Ionization potential. Atomic rays. Electronegativity. Electron affinity. Hydrides and oxides. Oxidation state.</p> <p>4. Chemical bonds, chemical compounds - structure and properties. Polarization of bonds. MO theory, symmetry and types of molecular orbits (LCAO). The theory of valence bonds (VB). Hybridization of orbitals and particle geometry. VSEPR method. Delocalized bindings. Metallic, hydrogen bonding, van der Waals. Bonding and geometric characteristics. Hydrogen. Characteristics and types of hydrides. Water, its physical and chemical characteristics. Association and dissociation of water. Construction of ice crystals. Hydrogen peroxide. Peroxides and suboxides. Peroxygen. Oxygen and its compounds.</p> <p>5. Acid-base reactions in solutions: Aqueous solutions. Electrolytes and non-electrolytes. Dissociation Electrolytic. Balance in electrolyte solutions. Constant and degree of electrolytic dissociation. Activity and activity factor. Ionic force. Product of solubility and activity. Acids, bases, salts. Theories: Arrhenius, Brønsted, Lewis. Balance. Amphotericism, hydrolysis, buffers, theory of indicators.</p> <p>Tutorials</p> <p>Basic concepts and chemical laws: Chemical substance, elements and chemical compounds. The law of mass preservation. The law of fixed and multiple relations. The law of simple volumetric relations. Atom and molecule. Mole. Atomic mass. Molecular atomic and molecular mass. Determination of experimental design and molecular formula. Gas law. The ideal gas status. Isothermal, isobaric and isochoric. Chemical compound and mixture. Solutions. Expressions of composition and concentration: percentage composition, mole fraction, units of concentration and their conversion (% , mol, ppm, ppb). Stoichiometry and chemical reactions: Chemical equation, balancing coefficients. Reminder of the nomenclature of basic inorganic compounds. Traditional names and systematic names of hydrides, oxides, hydroxides, acids and salts. Nomenclature of selected groups of organic compounds. Types of chemical reactions: synthesis, analysis, and exchange. The reactions occurring in solutions and precipitation reactions. Acids and rules. Neutralization reaction. Sol. Electrolytes and non-electrolytes. Electrolytes weak and strong. Acid-alkaline balance. Oxidation and reduction reactions.</p>											
Prerequisites and co-requisites	no requirements											
Assessment methods and criteria	<table border="1"> <thead> <tr> <th data-bbox="448 1688 794 1727">Subject passing criteria</th> <th data-bbox="794 1688 1141 1727">Passing threshold</th> <th data-bbox="1141 1688 1487 1727">Percentage of the final grade</th> </tr> </thead> <tbody> <tr> <td data-bbox="448 1727 794 1765">exercise: three written tests</td> <td data-bbox="794 1727 1141 1765">60.0%</td> <td data-bbox="1141 1727 1487 1765">40.0%</td> </tr> <tr> <td data-bbox="448 1765 794 1794">lecture: written exam</td> <td data-bbox="794 1765 1141 1794">60.0%</td> <td data-bbox="1141 1765 1487 1794">60.0%</td> </tr> </tbody> </table>			Subject passing criteria	Passing threshold	Percentage of the final grade	exercise: three written tests	60.0%	40.0%	lecture: written exam	60.0%	60.0%
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lecture: written exam	60.0%	60.0%										
Recommended reading	Basic literature	1. Bielański A., Podstawy chemii nieorganicznej. PWN, Warszawa, 2010 oraz wydania wcześniejsze. 2. Jones L., Atkins P.: Chemia ogólna. PWN, Warszawa, 2004 oraz wydania następne. 3. Cox P.A., Krótkie wykłady. Chemia Nieorganiczna, PWN, Warszawa, 2003. 4. KChNPG, skrypt on-line <a href="http://www.kchn.pg.gda.pl/?p=skrypt_cw">http://www.kchn.pg.gda.pl/?p=skrypt_cw</a>										

	Supplementary literature	1. Atkins P.: Podstawy chemii fizycznej. PWN, Warszawa, 2009.2. Sienko M., Plane R.: Chemia. Podstawy zastosowania. PWN, Warszawa, 1993.3. Pajdowski L.: Chemia ogólna. PWN, Warszawa, 1999.4. Kolditz L. (Ed.): Chemia nieorganiczna. PWN, Warszawa, 1994.5. Praca zbiorowa (Chmurzyński L., Gleich E., Myszka H., Nesterowicz M., Smiatacz K., Widernik T.): Obliczenia z chemii ogólnej. Wydawnictwo Uniwersytetu Gdańskiego, Gdańsk 2007
	eResources addresses	
Example issues/ example questions/ tasks being completed	1. What's allotropy? Give the allotropic varieties of oxygen. 2. Write the pressure constant of equilibrium for the Haber-Bosh ammonia reaction 3. Balance the given redox reaction 4. Write the electron configuration for atoms and ions: Cl, N, Na <sup>+</sup> , O <sup>2-</sup> 5. Draw formulas of Lewis SO <sub>2</sub> ; N <sub>3</sub> <sup>-</sup>	
Work placement	Not applicable	