



## Subject card

Subject name and code	Laboratory of the Organic Chemistry, PG_00037409						
Field of study	Biotechnology						
Date of commencement of studies	October 2020	Academic year of realisation of subject			2022/2023		
Education level	first-cycle studies	Subject group			Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	Full-time studies	Mode of delivery			at the university		
Year of study	3	Language of instruction			Polish		
Semester of study	5	ECTS credits			6.0		
Learning profile	general academic profile	Assessment form			assessment		
Conducting unit	Department of Organic Chemistry -> Faculty of Chemistry						
Name and surname of lecturer (lecturers)	Subject supervisor	dr hab. inż. Witold Przychodzeń					
	Teachers	dr hab. inż. Witold Przychodzeń dr hab. inż. Sebastian Demkowicz dr hab. inż. Teresa Olszewska dr inż. Monika Gensicka-Kowalewska dr hab. Sławomir Makowiec dr hab. Magdalena Śliwka-Kaszyńska dr inż. Jan Alfuth dr inż. Karol Biernacki					
Lesson types and methods of instruction	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	0.0	0.0	90.0	0.0	0.0	90
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan	Participation in consultation hours		Self-study	SUM	
	Number of study hours	90	6.0		54.0	150	
Subject objectives	Student should know, understand and use the basic methods and techniques used during the synthesis, isolation and purification of organic compounds						
Learning outcomes	Course outcome	Subject outcome			Method of verification		
	K6_W03				[SW1] Assessment of factual knowledge		
	K6_U03				[SU4] Assessment of ability to use methods and tools		
	K6_K06				[SK3] Assessment of ability to organize work		
Subject contents	Carrying out synthesis (one step or several steps) of selected organics belonging to different classes of compounds						
Prerequisites and co-requisites	Student must pass the classes and lectures in Organic Chemistry.						
Assessment methods and criteria	Subject passing criteria	Passing threshold			Percentage of the final grade		
	Collecting a sufficient amount of points (preparations, oral answers and laboratory introductory test).	60.0%			100.0%		

Recommended reading	Basic literature	1) K. Dzierzbicka, J. Rachoń, D. Witt - Preparatyka związków organicznych. Ćwiczenia laboratoryjne. 2) A. Vogel - Practical Organic Chemistry A. Vogel - Preparatyka organiczna. 3) March's Advanced Organic Chemistry. Reactions, mechanisms, and structure.
	Supplementary literature	1) J. Wróbel - Preparatyka i elementy syntezy organicznej. 2) M. Mąkosza - Synteza Organiczna. 3) B. Bochwic - Preparatyka Organiczna. 4) "Metabolic Basis of Detoxication. Metabolism of Functional Groups", Ed. W. B. Jakoby, AP, NY 1982.
	eResources addresses	Adresy na platformie eNauczenie:
Example issues/ example questions/ tasks being completed	<p>1. Calculate how many milliliters of water and concentrated hydrochloric acid (<math>d = 1.0 \text{ g/ml}</math>) you will use to obtain:</p> <p>a. 50 ml of 3 M HCl</p> <p>b. 18 ml of 12% HCl</p> <p>2. A mixture of 93 g benzyl alcohol and 62 g benzoic acid was heated to reflux in toluene. After 2 hours, 2.7 ml of water was collected in the Dean-Stark azeotropic apparatus. Draw the apparatus and calculate the yield of the reaction product assuming that the alcohol contained 6 mole% water and the acid purity was 90%.</p> <p>3. Draw the apparatus and give its application:</p> <p>a) reflux condenser with a spiral coil</p> <p>b) funnel with a pressure-equalising tube</p> <p>c) mercury manometer showing the pressure of 15 mm Hg</p> <p>d) Vigreux fractionating column with a swan-neck distillation adapter</p> <p>4. Describe in detail (giving the reaction schemes) the procedure of separation by extraction of a mixture consisting of 1,1-dimethoxyhexane, 2-aminophenol, phenylacetic aldehyde, N-methylbenzylamine and 2-nitrobenzoic acid into pure components (we start the process by dissolving equimolar amounts of the substance in ether and washing the organic phase with water).</p> <p>5. What color will the universal indicator paper color in the 0.1M aqueous solution (write the equation for the reaction):</p> <p>(a) boron tribromide (b) ammonium chloride (c) sodium acetate (d) n-butanol</p> <p>6. Draw a full steam distillation kit with a steam generator located on the left side of the kit.</p> <p>a) what melting point, volatility and solubility should the distilled substance have?</p> <p>b) what will be the consequences of clogging the condenser during distillation?</p> <p>c) calculate how many milliliters of water you will use to completely distill 10 g of xylene with steam? (temperature dependence on vapor pressure for xylene: <math>t [^{\circ}\text{C}] = 33.3 + 0.51 \text{ p [mm Hg]}</math>)</p> <p>7. Describe four main stages of crystallization from ethanol using activated carbon. Why is activated carbon ineffective when crystallized from hexane?</p> <p>8. Explain the terms: (a) equilibrium distillation, (b) pyrophoric substance, (c) positive azeotrope, (d) manostat, (e) Wood's alloy</p> <p>9. What funnel (slit-sieve or glass) is used to filter crystals after crystallization? Give at least three reasons for this choice</p> <p>10. Suggest a useful reagent for chemical neutralization of the following poisons. Write the equation of the reaction taking place.</p> <p>a) KCN b) <math>\text{NaN}_3</math> c) <math>\text{HgCl}_2</math></p> <p>1. Write the mechanism of the reaction of <math>(\text{CH}_3)_3\text{CCH}_2\text{Mg}</math> with 2,2,4-trimethylpentan-3-one leading to two C-8 products. 2. Explain the term "Schlenk equilibrium". How can we shift Schlenk equilibrium position? 3. What is the purpose of adding saturated ammonium chloride when processing crude 1,1-diphenylethanol?</p>	
Work placement	Not applicable	