



Subject card

Subject name and code	GENERAL CHEMISTRY, PG_00048911						
Field of study	Chemistry in Construction Engineering						
Date of commencement of studies	October 2023	Academic year of realisation of subject			2023/2024		
Education level	first-cycle studies	Subject group			Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	Full-time studies	Mode of delivery			at the university		
Year of study	1	Language of instruction			Polish		
Semester of study	1	ECTS credits			7.0		
Learning profile	general academic profile	Assessment form			exam		
Conducting unit	Department of Inorganic Chemistry -> Faculty of Chemistry						
Name and surname of lecturer (lecturers)	Subject supervisor	prof. dr hab. inż. Jarosław Chojnacki					
	Teachers						
Lesson types and methods of instruction	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	30.0	30.0	0.0	0.0	0.0	60
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	60		15.0		100.0	175
Subject objectives	Understanding of principles of general chemistry						
Learning outcomes	Course outcome		Subject outcome			Method of verification	
	K6_W03		The student has a well-established knowledge and good theoretical background in general chemistry, including the knowledge necessary to describe and understand the chemical phenomena and processes applied in construction engineering			[SW1] Assessment of factual knowledge	
	K6_U07		The student knows the basic theories about the structure of the atom and the molecules and about their mutual reactivity. He/she can use different ways of expressing concentrations of solutions and can calculate the amount of reagents necessary to obtain the desired substance by chemical reaction.			[SU3] Assessment of ability to use knowledge gained from the subject [SU4] Assessment of ability to use methods and tools	

Subject contents	<p>1. Structure of matter, the standard model.</p> <p>2. Electronic structure of the atom.</p> <p>3. Classification of the elements.</p> <p>4. Chemical bonds.</p> <p>5. Classification of chemical compounds.</p> <p>6. Chemical reactions.</p> <p>7. Concentrations of solutions.</p> <p>8. Chemical equilibria in water solutions.</p> <p>9. Basics of electrochemistry.</p> <p>10. Writing chemical reactions.</p> <p>11. Stoichiometric Calculations</p>											
Prerequisites and co-requisites												
Assessment methods and criteria	<table border="1"> <thead> <tr> <th data-bbox="456 595 794 629">Subject passing criteria</th> <th data-bbox="799 595 1137 629">Passing threshold</th> <th data-bbox="1142 595 1481 629">Percentage of the final grade</th> </tr> </thead> <tbody> <tr> <td data-bbox="456 636 794 669">Written tests for the classroom part</td> <td data-bbox="799 636 1137 669">53.0%</td> <td data-bbox="1142 636 1481 669">33.0%</td> </tr> <tr> <td data-bbox="456 676 794 701">Written exam for lectures</td> <td data-bbox="799 676 1137 701">55.0%</td> <td data-bbox="1142 676 1481 701">67.0%</td> </tr> </tbody> </table>			Subject passing criteria	Passing threshold	Percentage of the final grade	Written tests for the classroom part	53.0%	33.0%	Written exam for lectures	55.0%	67.0%
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Written exam for lectures	55.0%	67.0%										
Recommended reading	Basic literature	<p>1. L. Jones, P. Atkins, Chemia Ogólna. Cząsteczki, materia, reakcje. Wydawnictwo Naukowe PWN Warszawa 2014.</p> <p>2. A. Bielański, Podstawy Chemii Nieorganicznej, PWN Warszawa 2006</p> <p>3. Praca zbiorowa, Podstawy Obliczeń Chemicznych, Skrypt w wersji elektronicznej: https://chem.pg.edu.pl/kchn/chb-chemia-ogolna</p>										
	Supplementary literature	<p>1. M. J. Sienko, R. A. Plane, Chemia, Podstawy i Zastosowania, WNT 2002</p> <p>2. Z. Bądkowska, E. Koloński, M. Wojnowska, Obliczenia z Chemii Nieorganicznej, Wydawnictwo PG 1996 - skrypt.</p>										
	eResources addresses	Adresy na platformie eNauczanie:										
Example issues/ example questions/ tasks being completed	<p>Balance the reaction: $\text{MnO}_4^- + \text{SO}_3^{2-} + \dots = \text{Mn}^{2+} + \text{SO}_4^{2-} + \text{H}_2\text{O}$</p> <p>Give the electronic configuration of basic state and the number of unpaired electrons for Ga^+, N i F^-.</p> <p>Write chemical equations and name products of electrolysis of aqueous solution of CaCl_2 using platinum electrodes.</p>											
Work placement	Not applicable											