



Subject card

Subject name and code	Physical Chemistry , PG_00048440						
Field of study	Chemistry in Construction Engineering						
Date of commencement of studies	October 2023		Academic year of realisation of subject		2024/2025		
Education level	first-cycle studies		Subject group		Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	Full-time studies		Mode of delivery		at the university		
Year of study	2		Language of instruction		Polish		
Semester of study	3		ECTS credits		6.0		
Learning profile	general academic profile		Assessment form		exam		
Conducting unit	Department of Physical Chemistry -> Faculty of Chemistry						
Name and surname of lecturer (lecturers)	Subject supervisor		dr hab. inż. Maciej Śmiechowski				
	Teachers						
Lesson types and methods of instruction	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	30.0	30.0	30.0	0.0	0.0	90
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	90		5.0		55.0	150
Subject objectives	Cognition of physical laws governing chemical processes, thorough understanding of basic principles of thermodynamics allowing for effortless application of its conceptual framework in various disciplines of chemical sciences.						
Learning outcomes	Course outcome		Subject outcome		Method of verification		
	K6_W03		Student possesses a well-established and theoretically founded knowledge in the field of physical chemistry, including the knowledge necessary to describe and understand physicochemical phenomena and processes occurring in civil engineering as well as to measure and determine the parameters of these processes		[SW1] Assessment of factual knowledge		
	K6_K03		Student independently prepares reports on the physicochemical experiments, correctly estimating the measurement errors and confronting the obtained results with reliable literature values		[SK2] Assessment of progress of work [SK5] Assessment of ability to solve problems that arise in practice		
	K6_U07		Student independently solves the problems in basic thermodynamics, chemical equilibrium, phase equilibria and the basics of electrochemistry, using the known physicochemical laws		[SU3] Assessment of ability to use knowledge gained from the subject [SU1] Assessment of task fulfilment		
Subject contents	Basic concepts of phenomenological thermodynamics: the first and second law of thermodynamics and their consequences. Employing of thermodynamics in chemistry. The chemical equilibrium, Le Chatelier rule, dependence of equilibrium constant on temperature. Phase equilibria, the Clausius-Clapeyron equation, phase diagrams in a single and multi-component systems. Ideal and real solutions, activity coefficients. Principles of electrochemistry: the potential difference on the border of phases. Cells and electrode potentials. The polarization of electrodes. Surface phenomena and adsorption. Principles of chemical kinetics. Reaction rate, rate constant, order of reaction and activation energy, the influence of temperature on reaction rate. Catalysis.						

Prerequisites and co-requisites	Mathematics, Physics, General Chemistry, Technical Thermodynamics		
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
	Entry tests	60.0%	30.0%
	Final written exam	50.0%	40.0%
	Written colloquia	60.0%	30.0%
Recommended reading	Basic literature	<ol style="list-style-type: none">1. P. W. Atkins, Chemia fizyczna, Wydawnictwo Naukowe PWN, Warszawa 20012. P. W. Atkins, Podstawy chemii fizycznej, Wydawnictwo Naukowe PWN, Warszawa 19993. K. Pigoń, Z. Ruziewicz, Chemia fizyczna Tom 1. Podstawy fenomenologiczne, Wydawnictwo Naukowe PWN, Warszawa 20094. H. Strzelecki, W. Grzybowski (red.), Chemia fizyczna: ćwiczenia laboratoryjne, Wydawnictwo PG, Gdańsk 20045. I. Uruska (red.), Zbiór zadań z chemii fizycznej, Wydawnictwo PG, Gdańsk 1997	
	Supplementary literature	<ol style="list-style-type: none">1. H. Buchowski, W. Ufnalski, Podstawy termodynamiki, WNT 19942. A. Kiswa, Elektrochemia I. Jonika, WNT 2000.3. A. Kiswa, Elektrochemia II. Elektrodyka, WNT 20014. A. Molski, Wprowadzenie do kinetyki chemicznej, WNT 20015. M. R. Heal, A. R. Mount, A. G. Whittaker, Krótkie wykłady. Chemia fizyczna, Wydawnictwo Naukowe PWN, Warszawa 2003	
	eResources addresses	Adresy na platformie eNauczanie:	
Example issues/ example questions/ tasks being completed			
Work placement	Not applicable		