

Subject card

Subject name and code	Applied physical chemistry, PG 00066089								
Field of study	Chemia fizyczna stosowana								
Date of commencement of studies	February 2026		Academic year of realisation of subject			2025/2026			
Education level	second-cycle studies		Subject group			Obligatory subject group in the field of study			
						Subject group related to scientific research in the field of study			
Mode of study	Full-time studies		Mode of delivery			at the university			
Year of study	1		Language of instruction			Polish			
Semester of study	1		ECTS credits			3.0			
Learning profile	general academic profile		Assessment form			assessment			
Conducting unit	Department of Physic	al Chemistry -	> Faculty of Ch	emistry -> Fac	ulties of	Gdańs	k University o	f Technology	
Name and surname	Subject supervisor	dr hab. Aneta	r hab. Aneta Panuszko						
of lecturer (lecturers)	Teachers					- 			
Lesson types	Lesson type	Lecture 0.0	Tutorial 0.0	Laboratory 30.0	Project 15.0	:t	Seminar 0.0	SUM 45	
	Number of study hours	0.0	0.0	30.0	15.0		0.0	45	
	E-learning hours included: 0.0								
Learning activity and number of study hours	Learning activity	Participation in classes include plan		Participation in consultation hours		Self-study		SUM	
	Number of study hours	45		5.0		25.0		75	
Subject objectives	The aim of the course is to familiarize the student with various techniques, both experimental and theoretical, used to study the physicochemical properties of solutions and complex systems.								
Learning outcomes	Course out	Subject outcome			Method of verification				
	[K7_K02] is able to cooperate and work in a group, taking on different roles		The student effectively collaborates in a team, assuming various roles to accomplish common tasks and projects. The student takes responsibility for specific tasks and engages in communication and problemsolving within the group.			[SK5] Ocena umiejętności rozwiązywania problemów występujących w praktyce [SK1] Ocena umiejętności pracy w grupie [SK4] Ocena umiejętności komunikacji, w tym poprawności językowej			
	[K7_U02] prepares detailed documentation of the results of independently conducted experiments and analyzes the obtained results, uses professional vocabulary with understanding and prepares and communicates information		The student develops and interprets the results of independently conducted physicochemical experiments. The student presents the results in the form of a report.			[SU1] Ocena realizacji zadania [SU4] Ocena umiejętności korzystania z metod i narzędzi [SU5] Ocena umiejętności zaprezentowania wyników realizacji zadania			
	[K7_W03] recognizes and describes phenomena in the field of physics, including elements of quantum mechanics, solid state physics and nuclear physics, necessary to predict the course of physical phenomena and to solve technical problems		The student builds simple simulation systems and conducts their simulations using molecular dynamics methods. The student analyzes the results obtained through molecular dynamics methods and draws conclusions based on them.			[SW3] Ocena wiedzy zawartej w opracowaniu tekstowym i projektowym [SW1] Ocena wiedzy faktograficznej			
	[K7_U04] develops and transmits technical information in the form of text documents, spreadsheets, graphs, technological diagrams and multimedia presentations, and prepares a speech including a multimedia presentation		The student develops and delivers technical reports in the form of text documents and spreadsheets. The student creates charts and diagrams for result analysis. The student prepares a multimedia presentation and presents the results in the form of a public speech.			[SU3] Ocena umiejętności wykorzystania wiedzy uzyskanej w ramach przedmiotu [SU4] Ocena umiejętności korzystania z metod i narzędzi [SU5] Ocena umiejętności zaprezentowania wyników realizacji zadania			

Data wygenerowania: 06.11.2025 19:59 Strona 1 z 2

Subject contents	Course content – laboratory							
,	·							
	The subject of the course covers the following issues:							
	 Infrared spectroscopy in the study of intermolecular interactions: methods of developing spectra, selected chemometric methods in the analysis of spectral series in solving chemical problems. Analysis of the structure and energy of water hydrogen bonds around a solute: isotope dilution technique, difference spectra method, OD band spectral parameters, Badger-Bauer rule, empirical function correlating the oxygen-oxygen intermolecular distance with the position of the band. Intermolecular interactions in electrolyte solutions containing solvents with different donor properties and spatial structure and in aqueous solutions of non-electrolytes using precise measurements of density and sound speed: determining the limit values of apparent volumes and molar compressibility, division of standard partial volumes and molar compressibility of electrolytes into ionic shares, determination of solvation numbers of ions based on adiabatic compressibility and standard partial molar volume, coefficient of thermal expansion, Hepler's relation, standard partial volumes and molar compressivities of solute transfer, McMillan and Mayer equation. Heterogeneous redox reactions - calculation of the absolute value of the standard hydrogen electrode potential using the thermodynamic cycle. Standard potential of reduction reactions - determination, application and correlations with parameters reflecting the electronic structure of the molecule. Basics of molecular modeling methods: models of intermolecular interactions, taking into account constraints; problem of boundaries of simulation systems, periodic boundary conditions; molecular dynamics, numerical solution of equations of particle motion; different simulation conditions (NVT, NpT); results analysis methods: liquid structure, radial distribution functions, time correlation functions, diffusion. 							
Prerequisites	Mathematics, Physics, Physical Chemistry,							
and co-requisites								
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade					
	tests + reports (laboratory)	60.0%	50.0%					
	reports (project)	60.0%	50.0%					
Recommended reading	Basic literature	 H. Mark, J. Workman, Jr., "Chemometrics in spectroscopy", 2007, Elsevier J. Stangret, Zeszyty Naukowe Politechniki Gdańskiej, Chemia XLV (578), 2000. Organic Electrochemistry: Revised and Expanded (pp.229-259), Edition: 5th, Chapter: 6, Publisher: CRC Press, Taylor and Francis Group, Editors: Ole Hammerich, Bernd Speiser M.J. Abraham, D. van der Spoel, E. Lindahl, B. Hess, and the GROMACS Development Team. GROMACS 2023.4 Manual 						
	Supplementary literature	M. Śmiechowski, J. Stangret, Vibrational spectroscopy of semiheavy water (hdo) as a probe of solute hydration, Pure Appl. Chem. 82 (10) (2010) 18691887.						
	eResources addresses							
Example issues/ example questions/ tasks being completed	 What are the purposes of chemometric analysis methods for spectral data? Discuss the Badger-Bauer rule with a simple example. What fractions make up the standard partial molar volume of an electrolyte? Carrying out a computer simulation of argon gas under conditions of constant pressure and constant volume. 							
Practical activites within the subject	Not applicable							

Document generated electronically. Does not require a seal or signature.

Data wygenerowania: 06.11.2025 19:59 Strona 2 z 2