GDAŃSK UNIVERSITY OF TECHNOLOGY

Subject card

Subject name and code	INORGANIC CHEMISTRY, PG_00064378		
Field of study	Chemistry		
Date of commencement of studies	October 2026	Academic year of realisation of subject	2026/2027
Education level	first-cycle studies	Subject group	Obligatory subject group in the field of study Subject group related to scientific research in the field of study
Mode of study	Full-time studies	Mode of delivery	at the university
Year of study	1	Language of instruction	Polish
Semester of study	2	ECTS credits	7.0
Learning profile	general academic profile	Assessment form	assessment
Conducting unit	Department of Inorganic Chemistry -> Faculty of Chemistry -> Faculties of Gdańsk University of Technology		
Name and surname of lecturer (lecturers)	Subject supervisor	dr hab. inż. Łukasz Ponikiewski	
	Teachers		

Lesson types

Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
Number of study hours	30.0	15.0	60.0	0.0	0.0	105

E-learning hours included: 0.0

Learning activity and number of study hours

Learning activity	Participation in didactic classes included in study plan	Participation in consultation hours	Self-study	SUM
Number of study hours	105	5.0	65.0	175

Subject objectives

Types of chemical reactions - electron transfer, electron transfer, proton transfer reactions and ligand transfer reactions.Introduction of the students to the basic concepts of inorganic chemistry - properties of the elements and chemical compounds, their occurence in nature, processing and use. Part I. p-block elements

Learning outcomes

Course outcome	Subject outcome	Method of verification
[K6_W02] classifies acquired information, assessing its usefulness in solving the posed problems concerning the synthesis and analysis of selected groups of compounds, determining their physical and chemical properties, making measurements and determining the parameters of chemical reactions and processes	The student knows the chemical properties of p-block elements elements and their simple compounds. The student understands the connection between general chemistry rules and properties of simple chemical compounds.	[SW1] Assessment of factual knowledge [SW2] Assessment of knowledge contained in presentation
[K6_K03] is aware of the importance of caring for the quality and diligence of the tasks performed, being responsible for their consequences	The student is aware of the precision of the performed analyses and the consequences of incorrectly interpreting the information.	[SK2] Assessment of progress of work
[K6_U03] operates typical laboratory apparatus and carries out analyses to identify chemical compounds and materials, integrating computational methods and application software	The student knows how to perform basic qualitative analyses, characterize ions in aqueous solutions. In addition, he knows basic laboratory equipment.	[SU1] Assessment of task fulfilment [SU4] Assessment of ability to use methods and tools
[K6_U09] is able to recognise hazards, counteract them and work with chemical reagents and basic technical apparatus in accordance with health and safety principles and the concept of sustainability	The student knows the health and safety regulations necessary in working with chemical reagents. Is aware of the dangers of working with chemical reagents, especially concentrated acids and bases.	[SU3] Assessment of ability to use knowledge gained from the subject

Subject contents	Course content – lecture LECTURE: Redox- reactions. Acids and bases. The chemistry of nonmetals. Noble gases and their compounds. Halogens. The elements of groups 15 and 16 and their compounds with special emphasis on sulfur, nitrogen and phosphorus. The chemistry of group 14 elements - allotropes of carbon, inorganic compounds of carbon, silicon, silicates, silicones, germanium, tin and lead. Boron, boranes and oxoboranes. Course content – exercises EXERCISES: Solutions - solubility, concentrations percent, molar, normal, mol fraction, stoichiometry of the reactions in solutions. The concept of chemical equilibrium - basic calculations. Equilibria in the electrolyte solutions. Dissociation. Strong and weak electrolytes. The ion product of water. pH scale. Solutions of acids and bases. Solutions of salts. Buffer solutions. Precipitation equilibria and equilibria in solutions of complex compounds. Course content – laboratory LABORATORY: Every student has to do a two-semester course of classic qualitative analysis. During the running semester it consists of 7 practical exercises covering the qualitative analysis of selected cations.
Prerequisites and co-requisites	

Assessment methods and criteria

Subject passing criteria	Passing threshold	Percentage of the final grade
Lecture tests	60.0%	50.0%
Laboratory - short tests and detailed reports	45.0%	25.0%
Exercises - Two written tests during semester	60.0%	25.0%

Recommended reading		
	Basic literature	A. Bielański Chemia nieorganiczna, PWN recent editions; P.A. Cox Krótkie wykłady, chemia nieorganiczna, PWN 2003; F.A. Cotton, G. Wilkinson, P.L. Gaus Chemia nieorganiczna, podstawy, PWN 1995. University scripts: J. Prejzner: Inorganic Chemistry. Laboratory exercises. Issued by Gdansk University of Technology, Gdansk 2004. J. Chojnacki, A. Dołęga, S. Konieczny, A. Konitz, A. Okuniewski (red.), J. Pikies, A. Pladzyk, Ł. Ponikiewski, M. Walewski, A. Wiśniewska: Chemia ogólna i nieorganiczna. Ćwiczenia rachunkowe. *Wydawnictwo Politechniki Gdańskiej*, Gdańsk 2019. ISBN:978-83-7348-795-6.
	Supplementary literature	N.N. Greenwood, A. Earnshaw Chemistry of the elements Pergamon, 2nd Ed. (2005); C.E. Housecroft, A.G. Sharpe Inorganic chemistry, Pearson, Prentice Hall; 1st (2001), 2nd (2005) or 3rd (2008) editions; A.F. Wells Strukturalna chemia nieorganiczna WNT, 1993. M. Łaniecki Basics Inorganic Qualitative Analysis, Issued by UAM, Poznań; Calculations in General Chemistry, collective work, issued by University of Gdansk, Gdańsk.
	eResources addresses	

Example issues/ example questions/ tasks being completed

1) Why does the nitric oxide molecule have a permanent magnetic moment? Explain using the molecular orbitals diagram. Calculate the bond order in the nitric oxide molecule.

2) Why does iodine poorly dissolve in water and dissolve well in a solution of potassium iodide? Explain and write down the equation for the appropriate reaction.

3) List at least two carbon oxides, write down their names, draw Lewis formulas. Describe briefly the physical properties of these compounds (physical state, color, odor, solubility in water).

4) Write down the equations for the reactions of sodium chloride and sodium iodide with sulfuric acid (VI).

5) Describe the bonds found in the B2H6 molecule

6) Write down the reactions that occur in the production of nitric acid from ammonia. In which reaction is the use of a catalyst necessary? What kind of catalyst is used?

7) How is nitrogen obtained on a technical scale and how on a laboratory scale?

Practical activites within the subject	Not applicable

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